



MARK SCHEME for the May/June 2013 series

9701 CHEMISTRY

9701/21

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Cambridge is publishing the mark schemes for the May/June 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

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	GCE AS/A LEVEL – May/June 2013	9701	21

- 1 (a) (i) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (1)
- $(\text{NH}_4)_2\text{SO}_4 + 2\text{NaOH} \rightarrow 2\text{NH}_3 + \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$ (1)
- allow ionic equations in each case
- (ii) $n(\text{NaOH}) = n(\text{HCl}) = \frac{39.2 \times 2.00}{1000} = 0.0784$ (1)
- (iii) $n(\text{NaOH}) = n(\text{HCl}) = \frac{29.5 \times 2.00}{1000} = 0.059$ (1)
- (iv) $n(\text{NaOH}) = 0.0784 - 0.059 = 0.0194$ (1)
- (v) $n[(\text{NH}_4)_2\text{SO}_4] = \frac{0.0194}{2} = 9.7 \times 10^{-3}$ (1)
- (vi) mass of $(\text{NH}_4)_2\text{SO}_4 = 9.7 \times 10^{-3} \times 132.1 = 1.2814 \text{ g}$ (1)
- (vii) % of $(\text{NH}_4)_2\text{SO}_4 = \frac{1.2814 \times 100}{2.96} = 43.30405405 = 43.3$
 give one mark for the correct expression (1)
 give one mark for answer given as 43.3 – i.e. to 3 sig. fig. (1)
 allow ecf where appropriate [9]
- (b) fertiliser in the river causes
 excessive growth of aquatic plants/algae **or** algal bloom (1)
 when plants and algae die O_2 is used up **or** fish or aquatic life die (1) [2]
- (c) manufacture of HNO_3 **or** explosives **or** nylon **or**
 as a cleaning agent **or** as a refrigerant
not detergent (1) [1]

[Total:12]

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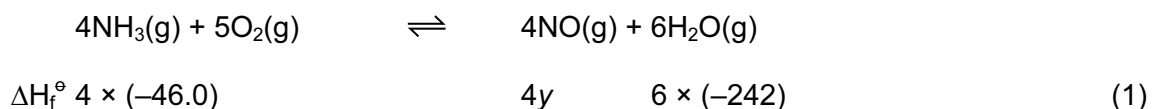
2 (a) $K_p = \frac{p(\text{NO})^4 p(\text{H}_2\text{O})^6}{p(\text{NH}_3)^4 p(\text{O}_2)^5}$ (1)

atmospheres **or** Pa **or** kPa (1)
allow ecf on incorrect powers [2]

(b) (i) **increasing temperature**
yield of NO is decreased **or** reaction moves to LHS (1)
forward reaction is exothermic (1)

(ii) **decreasing the pressure**
yield of NO is increased **or** reaction moves to RHS (1)
more moles/molecules of gas on RHS **or**
fewer moles/molecules of gas on LHS (1) [4]

(c) let ΔH_f° for NO be $y \text{ kJ mol}^{-1}$



$$\Delta H_{\text{reaction}}^\circ = 4y + [6 \times (-242)] - [4 \times (-46.0)] \qquad (1)$$

$$= 4y - 1452 + 184$$

$$\Delta H_{\text{reaction}}^\circ \text{ is } -906 \text{ kJ mol}^{-1} \text{ so} \qquad (1)$$

$$4y = -906 + 1452 - 184 = 362$$

$$\text{whence } y = \Delta H_f^\circ \text{ for NO} = +90.5 \text{ kJ mol}^{-1} \qquad (1)$$

+ sign is required [4]

[Total: 10]

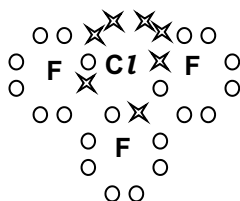
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3 (a) penalise (–1) for names of elements

- (i) Na **or** K **or** Li (1)
- (ii) S **or** C **or** N **or** P (1)
- (iii) K (1)
- (iv) C (1)
- (v) Cl (1)
- (vi) Al **or** Si (1) [6]

- (b) (i) Al₂O₃ **or** SiO₂ (1)
- (ii) Na₂O (1)
- (iii) P₂O₃ **or** P₄O₆ **and** P₂O₅ **or** P₄O₁₀ **or** SO₂ **and** SO₃ (1+1)
- (iv) Al₂O₃ (1) [5]

(c) (i)



- 3 bonding pairs **and**
 2 lone pairs around Cl atom (1)
 3 lone pairs on **each** of the F atoms (1)

(ii) **either**

referring to van der Waals' forces in BrF₃

- van der Waals' **or**
 intermolecular forces are greater/stronger (1)
 because there are more electrons in BrF₃ than in ClF₃ (1)

OR referring to permanent dipoles

- permanent dipole **or** intermolecular forces are stronger/greater in BrF₃ (1)
 because BrF₃ has a larger permanent dipole than ClF₃

OR because difference in electronegativity is larger between Br and F than between Cl and F (1)

part (ii) has a maximum of 2 marks (max 2) [4]

[Total: 15]

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4 Types of reaction used must come from the list in the question.

organic reaction	type of reaction	reagent(s)
CH ₃ CH ₂ CH ₂ CH ₂ Br → CH ₃ CH ₂ CH ₂ CH ₂ NH ₂	nucleophilic (1) substitution (1)	NH ₃ (1)
CH ₃ CH ₂ CH ₂ CH ₂ OH → BrCH ₂ CH ₂ CH ₂ CH ₂ OH	free radical (1) substitution (1)	Br ₂ or Br ₂ in an organic solvent (1) not Br ₂ (aq)
CH ₃ COCH ₃ → CH ₃ C(OH)(CN)CH ₃	nucleophilic (1) addition (1)	HCN or HCN and CN ⁻ or NaCN/KCN + H ⁺ (1)
CH ₃ CH(OH)CH ₂ CH ₃ → CH ₃ CH=CHCH ₃	elimination (1) not dehydration	conc. H ₂ SO ₄ or P ₄ O ₁₀ or Al ₂ O ₃ or H ₃ PO ₄ (1)

[Total: 11]

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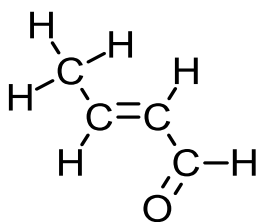
5 (a)

reaction	reagent	product
A	Br ₂ in an inert organic solvent	CH ₃ CHBrCHBrCHO
B	PCl ₃	NO REACTION
C	H ₂ and Ni catalyst	CH ₃ CH ₂ CH ₂ CH ₂ OH
D	NaBH ₄	CH ₃ CH=CHCH ₂ OH
E	K ₂ Cr ₂ O ₇ /H ⁺	CH ₃ CH=CHCO ₂ H

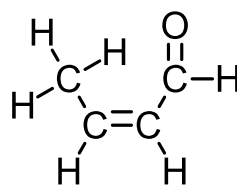
one mark for each correct answer

[5]

(b)



trans or E



cis or Z

two correct structures

(1)

both correctly labelled

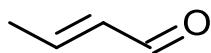
(1)

correctly displayed -CHO group

(1) [3]

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(c)



(1) [1]

(d) (i) $\text{CH}_3\text{CH}(\text{OH})\text{CH}(\text{OH})\text{CO}_2\text{H}$

(1)

(ii) $\text{CH}_3\text{CO}_2\text{H}$
 $\text{HO}_2\text{CCO}_2\text{H}$

(1)

(1) [3]

allow ecf on candidate's answer to E in (a)

[Total: 12]